



Mark Scheme (Results)

January 2023

Pearson Edexcel International Advanced Level
in Chemistry (WCH15)
Paper 01 Transition Metals and Organic Nitrogen
Chemistry

Question number	Answer	Mark
1	The only correct answer is D (Ni(CO)₄) <i>A is incorrect because copper has oxidation number +1 in [Cu(NH₃)₂]⁺ and nickel has oxidation number 0 in Ni(CO)₄</i> <i>B is incorrect because iron has oxidation number +3 in [Fe(CN)₆]³⁻</i> <i>C is incorrect because manganese has oxidation number + 2 in MnSO₄</i>	(1)

Question number	Answer	Mark
2(a)	The only correct answer is D (6) <i>A is incorrect because this is the number of types of ligand</i> <i>B is incorrect because this is the oxidation number of chromium</i> <i>C is incorrect because this is the number of ligands</i>	(1)

Question number	Answer	Mark
2(b)	The only correct answer is B (+1) <i>A is incorrect because there is a Cr³⁺ ion, two Cl⁻ ligands and the NH₂CH₂CH₂NH₂ ligands are neutral</i> <i>C is incorrect because there is a Cr³⁺ ion, two Cl⁻ ligands and the NH₂CH₂CH₂NH₂ ligands are neutral</i> <i>D is incorrect because there is a Cr³⁺ ion, two Cl⁻ ligands and the NH₂CH₂CH₂NH₂ ligands are neutral</i>	(1)

Question number	Answer	Mark
3	The only correct answer is A ($[\text{Al}(\text{H}_2\text{O})_6]^{3+}$) <i>B is incorrect because $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}$ gives a blue solution</i> <i>C is incorrect because $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$ gives an orange / brown solution</i> <i>D is incorrect because $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ gives a green solution</i>	(1)

Question number	Answer	Mark
4	The only correct answer is D ($\text{Zn}(\text{OH})_2$) <i>A is incorrect because $\text{Cu}(\text{OH})_2$ is only soluble in excess aqueous ammonia</i> <i>B is incorrect because $\text{Fe}(\text{OH})_2$ is insoluble in both excess aqueous ammonia and excess aqueous sodium hydroxide</i> <i>C is incorrect because $\text{Ni}(\text{OH})_2$ is only soluble in excess aqueous ammonia</i>	(1)

Question number	Answer	Mark
5	The only correct answer is C (24.4 %) <i>A is incorrect because 6.10 % is the value when only 1 mol of water is considered</i> <i>B is incorrect because 8.06 % is the value when only 1 mol of water is considered and no water has been included in the molar mass of the salt</i> <i>D is incorrect because 32.2 % is the value when no water has been included in the molar mass of the salt</i>	(1)

Question number	Answer	Mark
6	The only correct answer is C (diagram C with peak at (0.001, 8)) <i>A is incorrect because the complex ion with EDTA^{4-} has a more intense colour intensity than that with CN^- ions</i> <i>B is incorrect because EDTA^{4-} is a hexadentate ligand so the mol ratio Cr^{3+} ; EDTA^{4-} is 1 : 1 and the colour intensity should be higher</i> <i>D is incorrect because EDTA^{4-} is a hexadentate ligand so the mol ratio Cr^{3+} ; EDTA^{4-} is 1 : 1</i>	(1)

Question number	Answer	Mark
7	The only correct answer is D (activation energy is high) <i>A is incorrect because a positive value for E°_{cell} indicates the reaction is thermodynamically feasible</i> <i>B is incorrect because a positive value for $\Delta_r H$ would not be affected by a catalyst</i> <i>C is incorrect because a positive value for ΔS_{total} indicates the reaction is thermodynamically feasible</i>	(1)

Question number	Answer	Mark
8	The only correct answer is B ($\text{Fe}^{2+}(\text{aq}) \rightleftharpoons \text{Fe}^{3+}(\text{aq}) + \text{e}^-$ and $\text{Br}_2(\text{aq}) + 2\text{e}^- \rightleftharpoons 2\text{Br}^-(\text{aq})$) <i>A is incorrect because both half-equations cannot be oxidation</i> <i>C is incorrect because these half-equations would give $E^\circ_{\text{cell}} = -0.32 \text{ V}$</i> <i>D is incorrect because both half-equations cannot be reduction</i>	(1)

Question number	Answer	Mark
9	The only correct answer is C (525 (cm³)) <i>A is incorrect because 225 cm³ is the volume of oxygen needed to react with only 50 cm³ propene</i> <i>B is incorrect because 300 cm³ is the volume of oxygen needed to react with only 50 cm³ but-1-ene</i> <i>D is incorrect because 700 cm³ is the volume of oxygen needed to react with the whole mixture if the equations are balanced incorrectly by assuming that 1 mol of O₂ is needed to form 1 mol of H₂O</i>	(1)

Question number	Answer	Mark
10	The only correct answer is C (0.1 (cm³)) <i>A is incorrect because the volume of carbon dioxide is $250 \times 405/1 \times 10^6 = 0.101 \text{ cm}^3$, which is approximately 0.1 cm³</i> <i>B is incorrect because the volume of carbon dioxide is $250 \times 405/1 \times 10^6 = 0.101 \text{ cm}^3$, which is approximately 0.1 cm³</i> <i>D is incorrect because the volume of carbon dioxide is $250 \times 405/1 \times 10^6 = 0.101 \text{ cm}^3$, which is approximately 0.1 cm³</i>	(1)

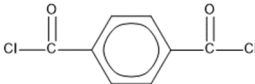
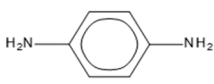
Question number	Answer	Mark
11(a)	The only correct answer is B (3095 – 3010 cm⁻¹) <i>A is incorrect because 3500 – 3300 cm⁻¹ shows the presence of N-H in an amine which is present in valine</i> <i>C is incorrect because 2962 – 2853 cm⁻¹ shows the presence of C-H in an alkane which is present in valine</i> <i>D is incorrect because 1725 – 1720 cm⁻¹ shows the presence of C=O in an carboxylic acid which is present in valine</i>	(1)

Question number	Answer	Mark
11(b)	The only correct answer is A (doublet and octet) <i>B is incorrect because the peak corresponding to the H attached to C with 2 CH₃ groups and CH will be an octet as there are 7 protons on the neighbouring carbons</i> <i>C is incorrect because the peak corresponding to the 2 CH₃ groups will be a doublet as there is 1 proton on the neighbouring carbon</i> <i>D is incorrect because the peak corresponding to the 2 CH₃ groups will be a doublet as there is 1 proton on the neighbouring carbon and the H attached to C with 2 CH₃ groups and CH will be an octet as there are 7 protons on the neighbouring carbons</i>	(1)

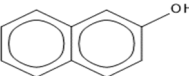
Question number	Answer	Mark
12(a)	The only correct answer is D (nucleophilic addition) <i>A is incorrect because electrophiles attack electron rich regions but the carbon atom attached to the magnesium is δ^-</i> <i>B is incorrect because the Grignard reagent is not a source of free radicals</i> <i>C is incorrect because increasing the length of the carbon chain is not oxidation</i>	(1)

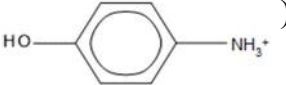
Question number	Answer	Mark
12(b)	The only correct answer is B (CH₃CH₂CH(CH₃)MgBr) <i>A is incorrect because this would form 1-phenylpentan-1-ol</i> <i>C is incorrect because this would form 1-phenyl-3-methylbutan-1-ol</i> <i>D is incorrect because this would form 1-phenyl-2,2-dimethylpropan-1-ol</i>	(1)

Question number	Answer	Mark
13	The only correct answer is C (0.6) <i>A is incorrect because the amino acid in the lower spot will have an R_f value of about 0.2</i> <i>B is incorrect because the distance moved by X is measured from the solvent front instead of from the baseline</i> <i>D is incorrect because the amino acid in the higher spot will have an R_f value of about 0.8</i>	(1)

Question number	Answer	Mark
14	The only correct answer is D ( and ) <i>A is incorrect because amines do not react with carboxylic acids</i> <i>B is incorrect because amides do react to form polyamides</i> <i>C is incorrect because this pair of monomers will not produce the required polyamide</i>	(1)

Question number	Answer	Mark
15(a)	The only correct answer is A (NaNO_2 and HCl at 5°C) <i>B is incorrect because NaNO_3 does not react with HCl to form the nitrous acid needed for the formation of benzenediazonium ions</i> <i>C is incorrect because nitrous acid and benzenediazonium ions decompose at 50°C</i> <i>D is incorrect because NaNO_3 does not react with HCl to form the nitrous acid needed for the formation of benzenediazonium ions and nitrous acid and benzenediazonium ions decompose at 50°C</i>	(1)

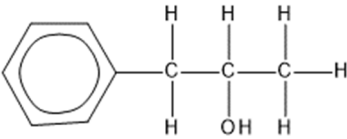
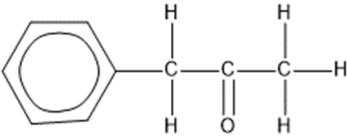
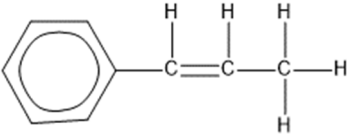
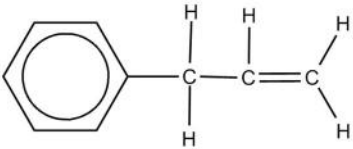
Question number	Answer	Mark
15(b)	The only correct answer is B ( in alkaline solution) <i>A is incorrect because alkaline conditions are needed to form an azo dye</i> <i>C is incorrect because the OH group is in the wrong position and alkaline conditions are needed to form an azo dye</i> <i>D is incorrect because the OH group is in the wrong position</i>	(1)

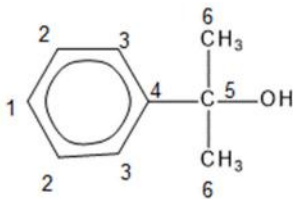
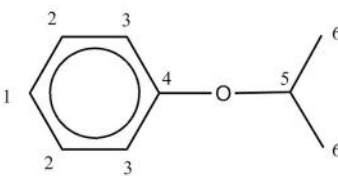
Question number	Answer	Mark
16	The only correct answer is A () <i>B is incorrect because the OH group will not be protonated in preference to the NH2 group</i> <i>C is incorrect because the addition of an acid will cause protonation not loss of a proton</i> <i>D is incorrect because the addition of an acid will cause protonation not loss of protons</i>	(1)

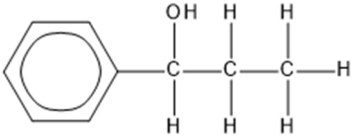
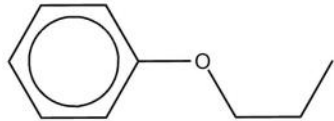
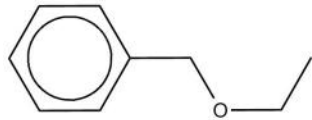
(Total for Section A = 20 marks)

Section B

Question Number	Answer	Additional Guidance	Mark
17(a)	- calculation of mol C / CO₂ (1) - calculation of mol H (1) - calculation of mol of O (1) - formula (1) Comment – alternative method via moles of A - M1 calculate moles of A : $5.26 \div 136 = 0.0387$ (mol) - M2 calculate moles of CO₂ and H₂O - M3 use combustion equation to show A forms 9CO₂ and 6H₂O, so C=9 and H=12 - M4 use of M_r to show mass due to O = 16, so number of O atoms = 1 If candidate does not score 4 marks, mark using method that gives best score	Example of calculation: mol CO₂ = $15.3/44 = 0.34773 = \text{mol C}$ or mass C = $15.3 \times 12/44 = 4.1727$ (g) and mol C = $4.1727/12 = 0.34773$ mol H₂O = $4.18/18 = 0.23222$ and mol H = $2 \times 0.23222 = 0.46444$ or mass H = $4.18 \times 2/18 = 0.46444$ (g) and mol H = 0.46444 mass O = $5.26 - 4.1727 - 0.46444 = 0.62286$ (g) and mol O = $0.62286/16 = 0.038929$ or moles of O = $5.26/136 = 0.03876$ (mol) TE on mass C and H ratio mol $\frac{0.34773}{0.038929}$ C : $\frac{0.46444}{0.038929}$ H : $\frac{0.038929}{0.038929}$ O 9 12 1 So formula is C₉H₁₂O (and this is the same as the molecular formula as $M_r = (9 \times 12) + (12 \times 1) + 16 = 136$) No TE on incorrect mol Ignore SF except 1 SF at each stage Comment : Ignore minor rounding errors e.g. 4.172 is acceptable for mass of C Allow masses of C, H and O to be determined and expressed as percentages e.g. C = 79.4%, H = 8.8%, so O = 11.8% Allow alternative methods	(4)

Question Number	Answer	Additional Guidance	Mark
17(b)	- structure of A (1) - structure of B (1) - structure of C (1) - structure of D (1) - structure of E (1)	Examples of structures Ignore connectivity of OH   CH₃ (standalone mark)  Allow <i>E</i> or <i>Z</i> isomer or both isomers  Allow displayed / structural formulae or any combination of these / skeletal formulae Allow Kekule Structures Allow D TE on incorrect 2° alcohol A e.g. if A is C₆H₅ CH(OH)C₂H₅ Allow E as TE on incorrect 1° alcohol A e.g. if A is C₆H₅CH₂CH₂CH₂OH If D and E are the wrong way around, allow (1) for M4 and M5	(5)

Question Number	Answer	Additional Guidance	Mark
17(c)	An explanation that makes reference to the following points: • structure of F • carbon atoms labelled	Example of structure:  Allow displayed / structural formulae or any combination of these / skeletal formulae Allow alternative clear ways of identifying carbon atoms Allow  M2 dependent on M1, unless very near miss (e.g. accidental omitting H on OH group)	(2)

Question Number	Answer	Additional Guidance	Mark
17(d)	An answer that makes reference to the following points: • structure of G (1) • m/z corresponds to $C_7H_7O^+$ / $C_6H_5CHOH^+$ / $C_6H_5OCH_2^+$ / $C_6H_5CH_2O^+$ / loss of C_2H_5 (1)	Example of structure  OR  OR  Allow any combination of structural and displayed formula / skeletal formula – if 2 structures are shown both must be incorrect Ignore missing + M1 and M2 are standalone marks No TE from incorrect structure	(2)

(Total for Question 17 = 13 marks)

Question Number	Answer	Additional Guidance	Mark
18(a)	+5 / 5+	Allow V / 5 / V ⁵⁺	(1)

Question Number	Answer	Additional Guidance	Mark
18(b)	An explanation that makes reference to the following points: - V⁵⁺ / the vanadium ion is (very) small / highly charged (1) - so it would polarise (two) water molecules / OH bonds (1) - causing them to lose hydrogen ions / H⁺ ions / deprotonate (1)	Allow high charge density Allow 'so it weakens OH bonds' Allow 'distorts electron clouds in water' Allow the energy required to remove 5 electrons to form V⁵⁺ is too high (1) because the energy is not recovered by the hydration of the ion (1) If no marks given allow 1 mark for correct electronic configuration of V⁵⁺ e.g. [Ar]	(3)

Question Number	Answer	Additional Guidance	Mark																																			
18(c)	- calculation of mol of each element (1) - deduction of mol ratio (and empirical formula) (1) - overall formula (1)	Example of calculation:	(3)																																			
		<table><tr><td></td><td>K</td><td>V</td><td>S</td><td>H</td><td>O</td></tr><tr><td>moles</td><td>$\frac{7.9}{39.1}$</td><td>$\frac{10.2}{50.9}$</td><td>$\frac{12.9}{32.1}$</td><td>$\frac{4.8}{1}$</td><td>$\frac{64.2}{16}$</td></tr><tr><td></td><td>= 0.202</td><td>= 0.200</td><td>= 0.402</td><td>= 4.8</td><td>= 4.01</td></tr><tr><td>divide by smallest</td><td>$\frac{0.202}{0.200}$</td><td>$\frac{0.200}{0.200}$</td><td>$\frac{0.402}{0.200}$</td><td>$\frac{4.8}{0.200}$</td><td>$\frac{4.01}{0.200}$</td></tr><tr><td></td><td>= 1</td><td>= 1</td><td>= 2</td><td>= 24</td><td>= 20</td></tr><tr><td colspan="5">Empirical formula KVS₂H₂₄O₂₀</td></tr></table>			K	V	S	H	O	moles	$\frac{7.9}{39.1}$	$\frac{10.2}{50.9}$	$\frac{12.9}{32.1}$	$\frac{4.8}{1}$	$\frac{64.2}{16}$		= 0.202	= 0.200	= 0.402	= 4.8	= 4.01	divide by smallest	$\frac{0.202}{0.200}$	$\frac{0.200}{0.200}$	$\frac{0.402}{0.200}$	$\frac{4.8}{0.200}$	$\frac{4.01}{0.200}$		= 1	= 1	= 2	= 24	= 20	Empirical formula KVS ₂ H ₂₄ O ₂₀				
				K	V	S	H	O																														
		moles		$\frac{7.9}{39.1}$	$\frac{10.2}{50.9}$	$\frac{12.9}{32.1}$	$\frac{4.8}{1}$	$\frac{64.2}{16}$																														
				= 0.202	= 0.200	= 0.402	= 4.8	= 4.01																														
		divide by smallest		$\frac{0.202}{0.200}$	$\frac{0.200}{0.200}$	$\frac{0.402}{0.200}$	$\frac{4.8}{0.200}$	$\frac{4.01}{0.200}$																														
				= 1	= 1	= 2	= 24	= 20																														
		Empirical formula KVS ₂ H ₂₄ O ₂₀																																				
		Ignore minor rounding errors in M1 and M2 TE on mol ratio from M1																																				
		Example of overall formula: KV(SO ₄) ₂ •12H ₂ O or K ₂ SO ₄ •V ₂ (SO ₄) ₃ •24H ₂ O																																				
Allow KV(SO ₄) ₂ (H ₂ O) ₁₂																																						
Allow K ₂ SO ₄ •V ₂ (SO ₄) ₃ (H ₂ O) ₂₄																																						
Ignore SF																																						
Allow the ions in any order / correct charges shown by individual ions, even if charges are not shown on all ions / missing dot(s) No TE from M2 to M3																																						

Question Number	Answer	Additional Guidance	Mark
18(d)(i)	- calculation of mol of V^{3+} (1) - calculation of mol MnO_4^- (1) - calculation of ratio of MnO_4^- to V^{3+} (1)	Example of calculation: $Mol V^{3+} = \frac{10.0 \times 0.132}{1000} = 0.00132 / 1.32 \times 10^{-3} (mol)$ $Mol MnO_4^- = \frac{26.40 \times 0.0200}{1000} = 0.000528 / 5.28 \times 10^{-4}(mol)$ Ratio $MnO_4^- : V^{3+} = 0.000528 : 0.00132$ $= 1 : 2.5$ or $2 : 5$ Allow TE on M1 and M2 Ignore SF except 1 SF in calculation of moles Allow 5 to 2 if it's clear it's $V^{3+} : MnO_4^-$	(3)

Question Number	Answer	Additional Guidance	Mark
18(d)(ii)	- correct species on each side of equation (1) - balancing (1) Comment – M2 dependent on M1	Example of equation: $2MnO_4^- + 5V^{3+} + 2H_2O \rightarrow 2Mn^{2+} + 5VO_2^+ + 4H^+$ Or $2MnO_4^- + 5V^{3+} + 22H_2O \rightarrow 2Mn^{2+} + 5[VO_2(H_2O)_4]^+ + 4H^+$ Allow multiples Allow (1) for $2MnO_4^- + 5V^{3+} + 16H^+ \rightarrow 2Mn^{2+} + 5V^{5+} + 8H_2O$ Ignore state symbols even if incorrect Allow oxidation to V(IV) if ratio 1 : 5 in (d)(i) $MnO_4^- + 5V^{3+} + H_2O \rightarrow Mn^{2+} + 5VO^{2+} + 2H^+$ species (1) balancing (1)	(2)

Question Number	Acceptable Answers	Additional Guidance	Mark												
18(e)*	This question assesses a student’s ability to show a coherent and logically structured answer with linkages and fully-sustained reasoning. Marks are awarded for indicative content and for how the answer is structured and shows lines of reasoning. The following table shows how the marks should be awarded for indicative content. <table><tr><td>Number of indicative marking points seen in answer</td><td>Number of marks awarded for indicative marking points</td></tr><tr><td>6</td><td>4</td></tr><tr><td>5–4</td><td>3</td></tr><tr><td>3–2</td><td>2</td></tr><tr><td>1</td><td>1</td></tr><tr><td>0</td><td>0</td></tr></table> The following table shows how the marks should be awarded for structure and lines of reasoning.	Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points	6	4	5–4	3	3–2	2	1	1	0	0	Guidance on how the mark scheme should be applied: The mark for indicative content should be added to the mark for lines of reasoning. For example, an answer with five indicative marking points that is partially structured with some linkages and lines of reasoning scores 4 marks (3 marks for indicative content and 1 mark for partial structure and some linkages and lines of reasoning). If there are no linkages between points, the same five indicative marking points would yield an overall score of 3 marks (3 marks for indicative content and no marks for linkages).	(6)
Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points														
6	4														
5–4	3														
3–2	2														
1	1														
0	0														

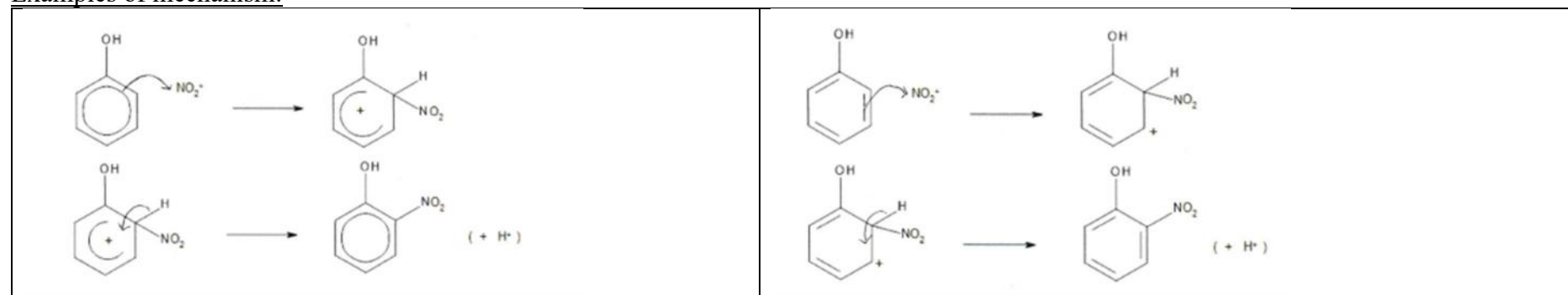
		Number of marks awarded for structure of answer and sustained line of reasoning	In general, it would be expected that 5 or 6 indicative points would get 2 reasoning marks, and 3 or 4 indicative points would get 1 mark for reasoning, and 0, 1 or 2 indicative points would score zero marks for reasoning. General points to note If there is any incorrect chemistry, deduct mark(s) from the reasoning. If no reasoning mark(s) awarded do not deduct mark(s).	
	Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.	2		
	Answer is partially structured with some linkages and lines of reasoning.	1		
	Answer has no linkages between points and is unstructured.	0		
	Comment: Look for the indicative marking points first, then consider the mark for structure of answer and sustained line of reasoning			

	Indicative content IP1 – vanadium(V) to vanadium(IV) Both iron and tin will reduce / convert / change V(V) to V(IV) and $E^\ominus_{\text{cell}}$ for Fe = (+)1.44 V and $E^\ominus_{\text{cell}}$ for Sn = (+)1.14 V IP2 – equations $2\text{VO}_2^+ + 4\text{H}^+ + \text{Fe} \rightarrow 2\text{VO}^{2+} + 2\text{H}_2\text{O} + \text{Fe}^{2+}$ and $2\text{VO}_2^+ + 4\text{H}^+ + \text{Sn} \rightarrow 2\text{VO}^{2+} + 2\text{H}_2\text{O} + \text{Sn}^{2+}$ IP3 – vanadium(IV) to vanadium(III) Both iron and tin will reduce / convert / change V(IV) to V(III) and $E^\ominus_{\text{cell}}$ for Fe = (+)0.78 V and $E^\ominus_{\text{cell}}$ for Sn = (+)0.48 V IP4 – equations $2\text{VO}^{2+} + 4\text{H}^+ + \text{Fe} \rightarrow 2\text{V}^{3+} + 2\text{H}_2\text{O} + \text{Fe}^{2+}$ and $2\text{VO}^{2+} + 4\text{H}^+ + \text{Sn} \rightarrow 2\text{V}^{3+} + 2\text{H}_2\text{O} + \text{Sn}^{2+}$ IP5 – vanadium(III) to vanadium(II) iron will reduce / convert / change V(III) to V(II) and $E^\ominus_{\text{cell}} = (+)0.18 \text{ V}$ and tin will not reduce / convert / change V(III) to V(II) and as $E^\ominus_{\text{cell}} = -0.12 \text{ V}$ IP6 – equation $2\text{V}^{3+} + \text{Fe} \rightarrow 2\text{V}^{2+} + \text{Fe}^{2+}$	Ignore state symbols in all equations even if incorrect If IP1 and IP2 not awarded, allow 1 IP for either totally correct iron or totally correct tin If IP3 and IP4 not awarded, allow 1 IP for either totally correct iron or totally correct tin Comment penalise references to Fe or Sn as oxidising agents once only in IP1, IP3 and IP5 Ignore any references to colour of vanadium species If no other marks awarded, allow 1 IP for idea that Fe can reduce to V^{2+} but Sn (only) to V^{3+} If no other marks awarded allow 1 IP for three pairs of correct $E^\ominus_{\text{cell}}$ values	
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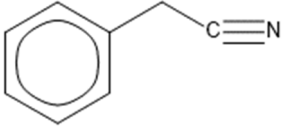
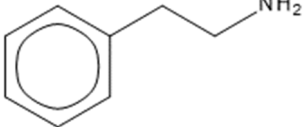
(Total for Question 18 = 18 marks)

Question Number	Answer	Additional Guidance	Mark
19(a)(i)	- curly arrow from on or within the circle to anywhere towards or on NO_2^+ (1) - intermediate structure including charge with horseshoe covering at least 3 carbon atoms and facing the tetrahedral carbon atom and some part of the positive charge must be within the horseshoe (1) - curly arrow from C-H bond to anywhere in the hexagon, and final organic product shown (1)	Allow arrow that starts from anywhere within the hexagon Do not award curly arrow starting on or outside the hexagon Do not award missing + on electrophile Do not award missing OH in M1 only Do not award dotted bonds to H and NO_2 unless they are part of a 3D structure Do not award formation of 4-nitrophenol / 3-nitrophenol in M2 only Comment – some part of the ‘horseshoe’ opening must be opposite the tetrahedral carbon, so only penalise if the line of the circle extends level with or past the tetrahedral C Ignore missing H^+ Ignore additional equations to generate NO_2^+ and reform catalysts	(3)

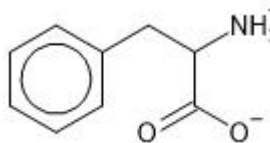
Examples of mechanism:

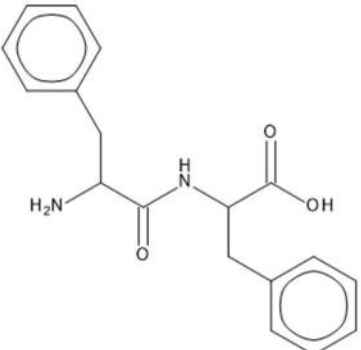


Question Number	Answer	Additional Guidance	Mark
19(a)(ii)	An explanation that makes reference to the following points: Phenol is more reactive than benzene / requires milder conditions because: - the lone pair (of electrons) on oxygen overlaps with the pi cloud / delocalised electrons / ring (1) - so increases the electron density of the (benzene) ring (1)	Allow reverse argument for why benzene is less reactive / requires harsher conditions Allow lone pair on OH group Ignore just lone pair Allow spreads into the pi cloud / delocalised electrons / ring (of electrons) Allow interacts with the pi cloud / delocalised electrons / ring (of electrons) Allow donated to the pi cloud / delocalised electrons / ring (of electrons) Allow the (benzene) ring is more susceptible to electrophilic attack Allow makes the (benzene) ring more nucleophilic Do not award 'makes the ring more electronegative'	(2)

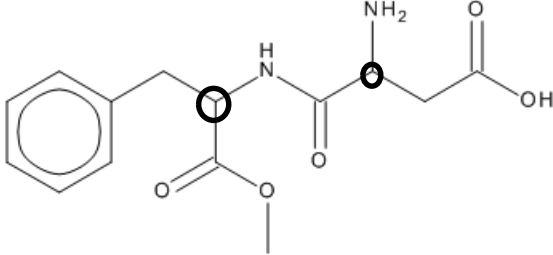
Question Number	Answer	Additional Guidance	Mark
19(b)	An answer that makes reference to the following points: - Reagent for step 1: potassium cyanide and (aqueous) ethanol (1) - First intermediate: 2-phenylethanenitrile (1) - Reagent for step 2: lithium tetrahydridoaluminate(III) / LiAlH_4 in (dry) ether (followed by (hydrolysis with) dilute acid / H^+) (1) - Second intermediate: 2-phenylethylamine (1) - Reagent for step 3: ethanoyl chloride / CH_3COCl / ethanoic anhydride / $(\text{CH}_3\text{CO})_2\text{O}$ (1)	Allow displayed / structural formulae or any combination of these / skeletal formulae for intermediates Ignore any references to heat / reflux throughout Ignore HCN Allow NaCN Stand alone mark  Allow -CN (i.e. triple bond not displayed) Do not award H_2 and Ni / Pt / Pd Stand alone mark  Allow TE from M2, if extra C shown in nitrile Do not award ethanoic acid / CH_3COOH Ignore AlCl_3	(5)

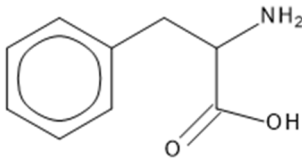
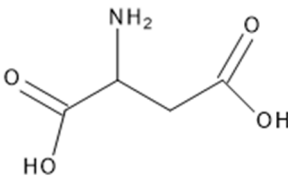
		Comment – allow (3) for use of Grignard reagent ; M1 –formation of Grignard, then reaction with HCHO to form $\text{C}_6\text{H}_5\text{CH}_2\text{CH}_2\text{OH}$, then reaction with KBr / H_2SO_4 to form $\text{C}_6\text{H}_5\text{CH}_2\text{CH}_2\text{Br}$, then reaction with NH_3 M2 structure of $\text{C}_6\text{H}_5\text{CH}_2\text{CH}_2\text{NH}_2$ M3 reaction of $\text{C}_6\text{H}_5\text{CH}_2\text{CH}_2\text{NH}_2$ with CH_3COCl	
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Question Number	Answer	Additional Guidance	Mark
19(c)(i)	structure of zwitterion	Example of zwitterion:  Allow '+' anywhere on NH₃ group Allow carboxylate ion shown with charge delocalised across two oxygen atoms Allow displayed / structural / skeletal formulae or any combination of these Ignore bond lengths and bond angles	(1)

Question Number	Answer	Additional Guidance	Mark
19(c)(ii)	structure of dipeptide	Example of dipeptide:  Allow displayed / structural / skeletal formulae or any combination of these Allow C₆H₅ for the phenyl groups Ignore bond lengths and bond angles Ignore connectivity of OH unless displayed as C-H-O (i.e. a bond shown from C to H then to O)	(1)

Question Number	Answer	Additional Guidance	Mark
19(c)(iii)	C₁₄H₁₈N₂O₅	Allow symbols in any order e.g. C ₁₄ H ₁₈ O ₅ N ₂	(1)

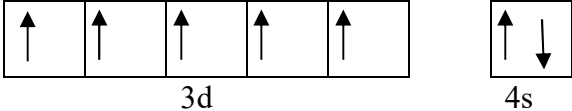
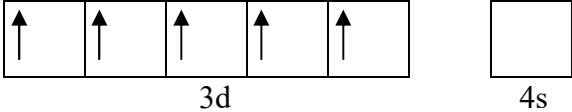
Question Number	Answer	Additional Guidance	Mark
19(c)(iv)	two chiral carbon atoms circled	Example of circled chiral carbons:  Allow other ways of representing the two carbon atoms e.g. asterisk * If more than two carbons are circled then do not award	(1)

Question Number	Answer	Additional Guidance	Mark
19(c)(v)	- structure of methanol (1) - structure of 2-amino-3-phenylpropanoic acid (1) - structure of 2-aminobutanedioic acid (1) Comment Ignore connectivity of OH unless displayed as C-H-O (i.e. a bond shown from C to H then to O)	Examples of structures: CH₃OH / —OH   Allow structures in any order Allow displayed / structural / skeletal formulae or any combination of these Accept NH₃⁺ for NH₂ groups Allow C₆H₅ for the phenyl group Ignore bond lengths and bond angles Ignore names even if incorrect	(3)

Question Number	Answer	Additional Guidance	Mark
19(c)(vi)	• calculation of mass of aspartame in 1 can (1) • calculation of mol of aspartame (1) • calculation of molecules of aspartame and answer to 1 / 2 / 3 SF (1) or • calculation of mol of aspartame in 100 cm³ (1) • calculation of mol of aspartame in 1 can (1) • calculation of molecules of aspartame and answer to 1 / 2 / 3 SF (1)	Examples of calculation: Mass of aspartame = $53 \times 330/100 = 174.9 \text{ mg} / 0.1749 \text{ g}$ Mol of aspartame = $0.1749/294 = 5.949 \times 10^{-4} / 0.0005949 \text{ (mol)}$ Molecules aspartame = $5.949 \times 10^{-4} \times 6.02 \times 10^{23}$ $(= 3.5813 \times 10^{20})$ $= 4 \times 10^{20} / 3.6 \times 10^{20} / 3.58 \times 10^{20}$ Or Mol aspartame = $53 \times 10^{-3}/294$ $= 1.8027 \times 10^{-4} / 0.00018027 \text{ (mol)}$ Mol aspartame in can = $1.8027 \times 10^{-4} \times 330/100$ $= 5.949 \times 10^{-4} / 0.0005949 \text{ (mol)}$ Molecules aspartame = $5.949 \times 10^{-4} \times 6.02 \times 10^{23}$ $(= 3.5813 \times 10^{20})$ $= 4 \times 10^{20} / 3.6 \times 10^{20} / 3.58 \times 10^{20}$ Allow alternative methods Allow TE throughout Correct answer with no working scores (3) Ignore SF except 1 SF in first 2 steps of working	(3)

(Total for Question 19 = 20 marks)
(Total for Section B = 51 marks)

Section C

Question Number	Answer	Additional Guidance	Mark
20(a)(i)	both electronic configurations correct	Examples of electronic configurations: Mn atom: [Ar]  Mn²⁺ ion: [Ar]  Allow half-arrows Allow all arrows pointing downwards in 3d subshell	(1)

Question Number	Answer	Additional Guidance	Mark
20(a)(ii)	An explanation that makes reference to the following points: • there is stability associated with a half-full set of (3)d orbitals (1) • Mn^{2+} has five d electrons so is more stable than Mn^{3+} (and has a higher E° value) (1) • Fe^{2+} has six d electrons so is less stable than Fe^{3+} (and has a lower E° value) (1) If M2 and M3 not awarded then allow 1 rescue mark for two correct electronic configuration from $\text{Mn}^{3+} = [\text{Ar}] 3d^4$ $\text{Fe}^{2+} = [\text{Ar}] 3d^6$ $\text{Fe}^{3+} = [\text{Ar}] 3d^5$ If more than two electronic configurations are given and one is incorrect then do not award the rescue mark	Allow reverse argument Allow 3d subshell with 5 electrons as alternative for half-filled Allow Mn^{2+} has five d electrons so eqm moves to RHS / Mn^{2+} has five d electrons so is energetically more favourable / more energy needed to remove an electron from Mn^{2+} as it has five d electrons Allow Fe^{3+} is more stable as it has a half-filled subshell so Fe^{2+} tends to lose electrons, (making E° less positive) Allow Fe^{2+} has a pair of electrons (in a d orbital) that repel so is less stable than Fe^{3+} (and has a lower E° value)	(3)

Question Number	Answer	Additional Guidance	Mark
20(b)(i)	An explanation that makes reference to the following points: - Mn reduced from (+)4 to (+)2 (1) - Cl oxidised from -1 to 0 and in Cl₂ (1)	Allow oxidation numbers shown under equation Allow Mn⁴⁺ and Mn²⁺ Allow Cl⁻ Comments: 0 must be linked to Cl₂ If no other mark is awarded, allow (1) for all oxidation numbers of Mn and Cl correct	(2)

Question Number	Answer	Additional Guidance	Mark
20(b)(ii)	- calculation of mol O₂ (1) - calculation of concentration of H₂O₂ (1)	Example of calculation: Mol of O₂ = $\frac{86.0}{24000} = 0.0035833 / 3.5833 \times 10^{-3} \text{ (mol)}$ Mol H₂O₂ = $2 \times 0.0035833 = 0.0071667 / 7.1667 \times 10^{-3} \text{ (mol)}$ and Conc H₂O₂ = $\frac{0.0071667 \times 1000}{100} = 0.071667 / 7.1667 \times 10^{-2} \text{ (mol dm}^{-3}\text{)}$ TE on mol O₂ Ignore SF except 1 SF Comment – if M1 is rounded to 0.00358 and carried through into M1 and M2, this gives a final answer of 0.0716	(2)

Question Number	Answer	Additional Guidance	Mark
20(c)	correct balanced equation	Example of equation: $3\text{MnO}_4^{2-} + 4\text{H}^+ \rightarrow 2\text{MnO}_4^- + \text{MnO}_2 + 2\text{H}_2\text{O}$ Allow multiples Allow reversible arrows provided reactants as shown are still on LHS Allow uncanceled electrons on either side Ignore state symbols even if incorrect Ignore oxidation states above atoms, even if incorrect	(1)

Question Number	Answer	Additional Guidance	Mark
20(d)	An explanation that makes reference to the following points: - this reaction is (auto)catalysed by the Mn^{2+} ions formed (1) (the reaction in) experiment 1 starts slowly because there is no Mn^{2+} / catalyst present initially (but speeds up as Mn^{2+} ions are formed) (1) (the reaction in) experiment 2 is fast(est) at the start as Mn^{2+} ions / catalyst (already) present (1)	Mn^{2+} can be mentioned at any point Allow experiment 1 starts slowly but speeds up as Mn^{2+} / catalyst forms Allow rate decreases constantly as Mn^{2+} ions / catalyst (already) present	(3)

Question Number	Answer	Additional Guidance	Mark
20(e)(i)	- calculation of mol of KMnO_4 (1) - calculation of mol of $\text{C}_6\text{H}_5\text{CH}_3$ (1) KMnO_4 is in excess because there are more than twice as many mol of KMnO_4 than mol of $\text{C}_6\text{H}_5\text{CH}_3$ (1)	Example of calculation: Mol $\text{KMnO}_4 = \frac{7.00}{158} = 0.044304 \text{ (mol)}$ Mol $\text{C}_6\text{H}_5\text{CH}_3 = \frac{1.73}{92} = 0.018804 \text{ (mol)}$ Accept 0.044304 mol KMnO_4 would react with 0.022152 mol $\text{C}_6\text{H}_5\text{CH}_3$ or reverse argument TE on M1 and M2 Allow other methods e.g. Mol $\text{C}_6\text{H}_5\text{CH}_3 = \frac{1.73}{92} = 0.018804 \text{ (mol) (1)}$ Minimum mass of KMnO_4 needed = $2 \times 0.018804 \times 158$ = 5.9421 (g) (1) This is less than 7 g so KMnO_4 is in excess (1) Ignore SF except 1 SF in M1 and M2	(3)

Question Number	Answer	Additional Guidance	Mark
20(e)(ii)	add H⁺ ions / acidify the solution / mixture	Allow correct name or formula of any strong acid e.g. HCl, H₂SO₄, HNO₃, H₃PO₄ Do not award carboxylic acids e.g. CH₃COOH Allow C₆H₅CO₂⁻ + H⁺ → C₆H₅COOH Ignore references to concentration / heat / reflux Do not award 'acid hydrolysis' / acid catalyst / H⁺ ions from water / inclusion of a second incorrect reagent e.g. H⁺ and LiAlH₄	(1)

Question Number	Answer	Additional Guidance	Mark
20(f)(i)	- anode half-equation (1) - cathode half-equation (1)	Examples of equations: $\text{Zn} + 2\text{OH}^- \rightarrow \text{ZnO} + \text{H}_2\text{O} + 2\text{e}^-$ $2\text{MnO}_2 + \text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{Mn}_2\text{O}_3 + 2\text{OH}^-$ Allow multiples / reversible arrows Ignore state symbols even if incorrect If no other mark is awarded allow (1) for anode and cathode half-equations written in wrong places If no other mark awarded allow 1 mark for Zn on the left-hand side of the anode reaction and MnO₂ on the left-hand side of the cathode reaction	(2)

Question Number	Answer	Additional Guidance	Mark
20(f)(ii)	$(E^\ominus = +)0.15 \text{ (V)}$	Do not award – 0.15(V)	(1)

(Total for Question 20 = 19 marks)
(Total for Section C = 19 marks)